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The video below refers to Dynamic Equilibrium (under the section Equilibrium Systems) 7.1 p. 420 in the Nelson Textbook The video below is related to 7.2 Equilibrium Law and the Equilibrium Constant (p.429 in the Textbook)

Pre University Courses

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Chemistry Notes for class 12 Chapter 2 Solutions

= 12.00 mmol - 5.40 mmol Unreacted ethanoic acid = 6.60 mmol Since 18.00 mL of NaOH(aq) was added to 20.00 mL of ethanoic acid solution, the total volume is now 38.00 mL. Solution: The concentration of unreacted ethanoic acid, [HC 2H 3O 2(aq)], can be determined: 23 2 23 2 HC H O (aq) 23 2 HC H O (aq) 23 2 [HC H O (aq)] (6.60 mmol) = 38.00 mL

Section 8.7: Acid-Base Titration Tutorial 1 Practice, page 547

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Gary S. Thorpe has taught AP Chemistry, College Chemistry, and gifted programs for over 30 years. Recipient of numerous awards in teaching. He is currently on staff at Beverly Hills High School, Beverly Hills, California. Author's Acknowledgments I would like to thank my wife, Patti, and my two daughters,

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