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What is a difference between gases and liquids. 1) Liquids have a stronger IMF. 2) Liquid particles are attracted to each other, whereas gas particles are not. What created the kinetic energy in liquids. the particles vibrating and spinning. what reduces the amount of space between the particles of a liquid.

Section 13.2 The Nature of Liquids Flashcards | Quizlet

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SECTION REVIEW THE NATURE OF LIQUIDS ANSWERS PDF

Class. Section Review. Objectives. Identify factors that determine physical properties of a liquid. Define evaporation in terms of kinetic energy. Describe the equilibrium between a liquid and its vapor. Identify the conditions under which boiling occurs. Vocabulary. vaporization.

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On this page you can read or download 13 3 the nature of solids section review answers in PDF format. If you don't see any interesting for you, use our search form on bottom ↓ . Session 2 The Particle Nature of Matter: Solids, Liquids, and

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SECTION 13.2 THE NATURE OF LIQUIDS (pages 390–395) This section describes a model for liquids in terms of kinetic energy and the attractive forces between the particles in a liquid. It also uses kinetic theory to distinguish evaporation from boiling. A Model for Liquids (page 390) 1. Is the following sentence true or false?

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Name Date Class STATES OF MATTER 13

The temperature at which a liquid changes into a solid. Crystal.

Solid in which the particles are arranged in an orderly way. ...

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14 Terms. madmik. 13.3 The Nature of Solids 22 Terms. eaecco.

Science chapter 2, lesson 1 book info (solids) 17 Terms.

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the normal boiling point of a liquid is the temperature at which the vapor pressure is equal to ____ 101.3 kPa/1 atm/760 mmHg

T/F: particles in a liquid don't have enough kinetic energy to overcome the attractive forces between them and vaporize them

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Start studying 10.2 The Nature of Liquids. Learn vocabulary, terms, and more with flashcards, games, and other study tools.

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A measure of the force exerted by a gas above a liquid in a sealed container; a dynamic equilibrium exists between the vapor and the liquid. Boiling Point The temperature at which the vapor pressure of a liquid is just equal to the external pressure on the liquid.

13.2 The Nature of Liquids Flashcards | Quizlet

Section 2.1 Liquids, Solids, and Gases Goals To describe a model that allows you to visualize the particle nature of matter. To describe the similarities and differences among solids, liquids, and gases in terms of this model. This is a very important section because it presents a model that you will use throughout your

Chapter 2 The Structure of Matter and the Chemical Elements

Liquids Class Date 13.1 Section Review DIRECTIONS: Write on the line at the right of each statement the letter preceding the word or expression that best completes the statement

Compared to the particles in a gas, the particles in a liquid (a) have more energy;

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An editor will review the submission and either publish your submission or provide feedback. Next Answer Chapter 13 - States of Matter - 13.4 Changes of State - Chemistry & You - Page 439: Q Previous Answer Chapter 13 - States of Matter - 13.3 The Nature of Solids - 13.3 Lesson Check - Page 434: 23

Chapter 13 - States of Matter - 13.3 The Nature of Solids

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A Model for Liquids According to the kinetic theory, both the particles in gases and the particles in liquids have kinetic energy. This energy allows the particles in gases and liquids to flow past one another, as shown in Figure 13.5. Substances that can flow are referred to as fluids.

13.2 The Nature of Liquids - Henry County School District

Section 3.1 Liquids, Solids, and Gases Goals To describe a model that allows you to visualize the particle nature of matter. To describe the similarities and differences among solids, liquids, and gases in terms of this model. This is a very important section because it presents a model that you will use throughout your

Chapter 3 The Structure of Matter and the Chemical Elements

solubilities of liquids or solids o Increase in pressure increases the solubility of gases Henry's Law o Solubility of a gas in a liquid is directly proportional to the partial pressure of that gas on the surface of the liquid n ex. pop n When pressure is released gas escapes as bubbles

CHEMISTRY - Chapter 13

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